

Name: _____ Date: _____

Molarity and Dilution Practice

Molarity Problems: Solve the following molarity problems. Round all answers to the correct number of significant figures. Write out all answers to the correct unit.

- What is the molarity of a 50.0 mL solution that contains 0.125 moles of NaCl ?

$$M = \frac{n}{V} = \frac{0.125 \text{ mol}}{0.0500 \text{ L}} = 2.50 \text{ M}$$
- What is the molarity of a solution that contains 0.025 moles of sodium hydroxide, NaOH , in 100 mL of solution? The molar mass of sodium hydroxide is 40.0 g/mol.

$$M = \frac{n}{V} = \frac{0.025 \text{ mol}}{0.100 \text{ L}} = 0.250 \text{ M}$$
- What mass of zinc nitrate, $\text{Zn}(\text{NO}_3)_2$, is needed to prepare 100 mL of a 0.100 M zinc nitrate solution?

$$n = M \times V = 0.100 \text{ M} \times 0.100 \text{ L} = 0.0100 \text{ mol}$$

$$m = n \times M_r = 0.0100 \text{ mol} \times 297 \text{ g/mol} = 2.97 \text{ g}$$

Dilution Problems: Solve the following dilution problems. Round all answers to the correct number of significant figures. Write out all answers to the correct unit.

- Substance A is a dilute solution of a concentration of 0.100 M. A 100 mL volume of this solution is added to enough water to make a total volume of 250 mL.

$$M_1 V_1 = M_2 V_2$$

$$0.100 \text{ M} \times 100 \text{ mL} = M_2 \times 250 \text{ mL}$$

$$M_2 = \frac{0.100 \text{ M} \times 100 \text{ mL}}{250 \text{ mL}} = 0.0400 \text{ M}$$
- How many milliliters of 0.100 M sodium hydroxide must be diluted with water to prepare 100 mL of a 0.0250 M sodium hydroxide solution?

$$M_1 V_1 = M_2 V_2$$

$$0.100 \text{ M} \times V_1 = 0.0250 \text{ M} \times 100 \text{ mL}$$

$$V_1 = \frac{0.0250 \text{ M} \times 100 \text{ mL}}{0.100 \text{ M}} = 25.0 \text{ mL}$$

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Molarity By Dilution Worksheet Answers Hudson City Schools / Homepage Created Date: 1/8/2016 9:43:18 AM Hortonville Area School District | Our community ensures ... Created Date: 3/18/2015 4:06:29 PM Molarity Problems Worksheet - Mrs Getson's Blog Dilution Problems Worksheet 1 How do you prepare a 250-ml of a 235 M HF dilution from a 150 M stock solution? 2 If 455-ml of 60 M HNO₃ is used to ...

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Molarity 1 (Worksheet) Last updated; Save as PDF Page ID 90889; No headers. Name: _____ Section: _____ Student ID#: _____ Work in groups on these problems. You should try to answer the questions without referring to your textbook. If you get stuck, try asking another group for help. Calculate molarity if 25.0 mL of 1.75 M HCl diluted to 65.0 mL. Calculate molarity by dissolving 25.0g NaOH in ...

Molarity 1 (Worksheet) - Chemistry LibreTexts

answers, Chemistry dilution practice, Dilutions work name key, Solutions work 2 molarity and dilution problems answers. Dilution Worksheets - Kiddy Math Dilutions Worksheet - Solutions 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be?

Molarity And Dilution Worksheet Answers

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Molarity And Dilution Worksheet Answers

Dilutions Worksheet 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted

solution be? 3) How much 0.05 M HCl solution can be made by diluting 250 mL of 10 M HCl? 4) I have 345 mL of a 1.5 M NaCl solution. If ...

Dilutions Worksheet - Awesome Science Teacher Resources

3. What is the molarity of a solution of HNO₃ that contains 12.6 grams HNO₃ in 1.0 L of solution? ? mol HNO₃ = 12.6 g HNO₃ × 1 mol HNO₃ / 63.0 g HNO₃ = 0.200 mol HNO₃ M = 0.200 mol HNO₃ / 1.0 L = 0.200 M 4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution? 0.700 M = moles of solute / 0.250 L moles of ...

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Dilution Worksheets - Kiddy Math

Dilution Problems Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 39.2 mL. 2. If 455-ml of 6.0 M HNO₃ is used to make a 2.5 L dilution, what is the molarity of the dilution? 1.1 M. 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is the molarity of the ...

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