

Calculate the molarity of a solution prepared by dissolving 1.56 grams of gaseous HCl in enough water to make 26.8 mL of solution.

$$1.56 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.46 \text{ g}} = 4.28 \times 10^{-2} \text{ mol}$$

$$26.8 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 2.68 \times 10^{-2} \text{ L}$$

$$M = \frac{4.28 \times 10^{-2} \text{ mol}}{2.68 \times 10^{-2} \text{ L}} = 1.6$$

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MOLARITY BY DILUTION Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid $M \times V$ then, $M_1 \times V_1 = M_2 \times V_2$. Solve the following problems. 1. How much concentrated 18 M sulfuric acid is needed to prepare 250 ml- of a 6.0 M ...

Newbury Park High School

Solve the Dilution: What is the Molarity? 350.mL of 12 M HCl is dissolved into 850. mL of water. Question. Asked Sep 24, 2020. 1 views. Solve the Dilution: What is the Molarity? 350.mL of 12 M HCl is dissolved into 850. mL of water; check circle Expert Answer. Want to see the step-by-step answer? See Answer. Check out a sample Q&A here. Want to see this answer and more? Step-by-step answers ...

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