

Name _____ Date _____ Class _____

12 STOICHIOMETRY

Chapter Test B

A. Matching

Match each item in Column A with the correct description in Column B. Write the letter of the correct item on the line.

Column A	Column B
_____ 1. An addition that describes the amount of product that can be formed in a reaction	a. percent yield
_____ 2. The amount of product that forms when a reaction is carried out in the laboratory	b. limiting reagent
_____ 3. An calculation of quantities in chemical equations	c. theoretical yield
_____ 4. The ratio of the actual yield to the theoretical yield expressed as a percent	d. stoichiometry
_____ 5. An addition that is given to enough quantity to react with a limiting reagent	e. actual yield
_____ 6. The maximum amount of product that could be formed from given amounts of reactants	f. excess reagent

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 7. In a chemical reaction
- | | |
|------------------------|--------------------------------------|
| a. mass is conserved | c. matter is conserved |
| b. atoms are conserved | d. both mass and atoms are conserved |
- _____ 8. Which of the following is a correct decomposition of the balanced equation?
- $$2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$$
- | | |
|--|--|
| a. 2 grams H_2O = 2 grams H_2 + 1 gram O_2 | c. 2 moles H_2O = 2 moles H_2 + 1 mole O_2 |
| b. 2 grams H_2 + 2 grams H_2O = 4 grams H_2O + 2 grams H_2 | d. both a and c |
| e. 2 moles H_2 + 2 moles H_2O = 4 moles H_2O + 2 moles H_2 | |
- _____ 9. If 10 moles of HCl are consumed in the reaction below, how many moles of H_2 are produced?
- $$2\text{HCl} + \text{Fe} \rightarrow \text{H}_2 + \text{FeCl}_2$$
- | | |
|-----------|-----------|
| a. 10 mol | c. 20 mol |
| b. 5 mol | d. 40 mol |

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Chapter Test B Stoichiometry Answers

Chapter Test B Chapter: Stoichiometry PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. Knowing the mole ratio of a reactant and product in a chemical reaction would allow you to determine a. the energy released in the reaction. b. the speed of the reaction.

Name Class Date Assessment) Chapter Test B

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Stoichiometry Chapter Test B

Modern Chemistry 69 Chapter Test Chapter: Chemical Equations and Reactions PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The production of a slightly soluble solid compound in a double-displacement reaction results in the formation of a a. gas. b ...

Assessment Chapter Test B - Ed W. Clark High School

Remember it is a MC test, use the answers • Mark which questions you would like to “go over” when we get to school in September. 1. Balance the following equation: ____NH₃ + ____O₂ → ____NO₂ + ____H₂O The balanced equation shows that 1.00 mole of NH₃ requires ____ mole(s) of O₂ a. 0.57 b. 1.25 c. 1.33 d. 1.75 e. 3.5 2. Write a balanced equation for the combustion of acetaldehyde, CH₃CHO ...

Practice Test Ch 3 Stoichiometry Name Per

CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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