

20. Calculate the pH at the equivalence point after the titration of each of the following with 0.100 M NaOH:

( $K_a$  for acetic acid is  $1.8 \times 10^{-5}$ )

a) 1.0 L of 0.100 M HBr

pH = 7.00 b/c s.a. vs. s.b.

b) 1.0 L of 0.100 M HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> → C<sub>2</sub>H<sub>3</sub>O<sub>2</sub><sup>-</sup> w. base  $\sqrt{K_b \cdot [ ]}$   
 $\text{mol} = M \cdot L = 1.0(0.1) = 0.1 \text{ mol}$   
 $\frac{0.1 \text{ mol}}{2 \text{ L}} = \boxed{M}$

21. How many moles of sodium hypobromite (NaBrO) should be added to 1.00 L of 0.056 M hypobromous acid (HBrO) to  
 ScreenstOfAnt.com pH 9.29? ( $K_a$  for HBrO =  $2.5 \times 10^{-9}$ )

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## Chapter 16 Ap Chemistry Answers

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### Chapter 16. Acid-Base Equilibria

Chapter 16; Chapter 17; Chapter 18; Chapter 19; Chapter 20; Chapter 21; Index; 1. A reaction has a natural tendency to occur and takes place without the continual input of energy from an external source. 3. (a) spontaneous; (b) nonspontaneous; (c) spontaneous; (d) nonspontaneous; (e) spontaneous; (f) spontaneous. 5. Although the oxidation of plastics is spontaneous, the rate of oxidation is ...

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AP Chemistry Chapter 15 Answers - Zumdahl 15.65 HIn = In- + H + K a = [In-][H +] / [HIn] = 1.0 x 10-9 a. In a very acid solution, the HIn form dominates, so the solution will be yellow. b. The color change occurs when the concentration of the more dominant form is approximately ten times as great as the less dominant form of the indicator.

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