

## Chemistry Chapter 10-12 Review Worksheet Answer Key (except multiple choice)

## Section 10.1- The Mole: A Measurement of Matter

1. a. 208.2 g/mol    b. 352.0 g/mol    c. 158.0 g/mol    d. 310.2 g/mol  
 2. 5.85 mol H<sub>2</sub>O    3. 3.6 X 10<sup>23</sup> atoms    4. 32.0 g

## Section 10.2- Mole-Mass and Mole-Volume Relationships

1. a. 1.8 x 10<sup>3</sup> g    b. 3.20 x 10<sup>-2</sup> g    c. 0.480 g    d. 1.43 x 10<sup>2</sup> g  
 2. 1.87 X 10<sup>2</sup> g  
 3. a. 4.9 X 10<sup>-3</sup> mol    b. 9.10 X 10<sup>-7</sup> mol    c. 1.98 X 10<sup>-5</sup> mol    d. 1.98 X 10<sup>-5</sup> mol  
 4. 15.1 g  
 5. 59.6 L CH<sub>4</sub>

## Section 10.3- Percent Composition and Chemical Formulas

1. Percent C = (5.34 g C / 52.84 g) x 100 = 10.1% C  
 Percent H = (0.42 g H / 52.48 g) x 100 = 0.79%  
 Percent Cl = (47.08 g Cl / 52.84 g) x 100 = 89.1% Cl  
 2. Mass of Cl = 18.35 - 5.74 g = 12.61 g Cl  
 % Sn = (5.74 g Sn / 18.35 g compound) x 100 = 31.3% Sn  
 3. Percent C = (48.0 g C / 158.1 g Ca(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)) x 100 = 30.4% C  
 0.304 x 65.3 = 19.8 g C  
 4. 13.2 g

## Section 11.1- Describing Chemical Reactions

1. Fe(s) + S(s) → FeS(s)  
 2. MgCO<sub>3</sub>(s) → MgO(s) + CO<sub>2</sub>(g)  
 3. silver + sulfur → silver sulfide  
 Silver metal and sulfur react to produce solid silver sulfide. There are 2 silver atoms and 1 sulfur atom on each side of the equation, and the coefficients are in their lowest possible ratio. Thus, the equation is balanced correctly.  
 4. 2AgNO<sub>3</sub>(aq) + Cu(s) → 2Ag(s) + Cu(NO<sub>3</sub>)<sub>2</sub>(aq)

## Section 11.2- Types of chemical Reactions

1. 2Al(s) + 3F<sub>2</sub>(g) → 2AlF<sub>3</sub>(s)  
 2. Ca(s) + 2HCl(aq) → H<sub>2</sub>(g) + CaCl<sub>2</sub>(aq)  
 3. C<sub>3</sub>H<sub>8</sub>(g) + 5O<sub>2</sub>(g) → 3CO<sub>2</sub>(g) + 4H<sub>2</sub>O(g)  
 4. FeCl<sub>3</sub>(aq) + 3NaOH(aq) → Fe(OH)<sub>3</sub>(s) + 3NaCl(aq)  
 5. combination reaction: 1    single-replacement reaction: 2  
 double-replacement reaction: 4    combustion reaction: 3  
 6. a. no reaction  
 b. Ca(s) + Mg(NO<sub>3</sub>)<sub>2</sub>(aq) → Ca(NO<sub>3</sub>)<sub>2</sub>(aq) + Mg(s)

## Section 12.1 - The Arithmetic of Equations

1. 10A + 2C + Ci → A<sub>10</sub>C<sub>2</sub>Ci  
 25 A<sub>10</sub>C<sub>2</sub>Ci x 10A / A<sub>10</sub>C<sub>2</sub>Ci = 250 apples

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country, but those are fairly minor quibbles. Chapter 11 Chemical Reactions Vocabulary Chapter 11 Chemical Reactions Vocabulary. STUDY. PLAY. Product. A new substance formed in a ...

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Chapter 11 "Describing Chemical Reactions" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to study the vocabulary and major concepts presented in this chapter. A B; combustion reaction: the complete reaction of a hydrocarbon with oxygen to produce carbon dioxide and water: decomposition reaction : a single compound is broken ...

### **Quia - Chapter 11 "Describing Chemical Reactions"**

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### **Chapter 11 Chemical Reactions Vocabulary Review Answer Key**

Chapter 11 Chemical Reactions Crossword Down : 1) A chemical change that involves an exchange of positive ions between two compounds. 2) A chemical equation that does not indicate the relative amounts of the reactants and products.

### **Vocabulary Crossword Puzzle: Chapter 11 Chemical Reactions ...**

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### **i½i½' [Book] Chapter 11 Chemical Reactions Vocabulary ...**

Chapter 11 "Describing Chemical Reactions" Use this activity to test your knowledge of the vocabulary and major concepts presented in this chapter. Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Print; Help; Mrs. Athens. Evangelical Christian School. View profile ; This activity was created by a Quia Web subscriber. Learn more about Quia: Create your own ...

### **Quia - Chapter 11 "Describing Chemical Reactions"**

330 Chapter 11 11.2 Types of Chemical Reactions Often charcoal briquettes provide the heat for barbecue grills through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal grill? The heat and smoke are the products of a combustion reaction. Combustion is one of the five general types of

chemical

### **11.2 Types of Chemical Reactions - Henry County School ...**

• Predict the products of the five general types of reactions Vocabulary Part ACompletion Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number. It is possible to the products of some chemical 1. reactions. In order to do this, you must be able to recognize at ...

### **05 CTR ch11 7/9/04 3:33 PM Page 265 DESCRIBING CHEMICAL ...**

368 Chapter 11 • Stoichiometry Section 11.11.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry MAIN Idea The amount of each reactant present at the start of ...

### **Chapter 11: Stoichiometry - Mr. Miller's Classes**

In a chemical reaction, one or more reactants change into one or more products. Figure 11.1a shows the ingredients for making leavened bread— flour, salt, yeast, and water. Chemical reactions take place when the ingre- dients are mixed together and heated in the oven. Figure 11.1b shows the product—a loaf of bread.

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