

Calculating pH
The pH of a solution is a logarithmic relationship.

What

Calculate the pH of a solution of 0.10 M acetic acid (CH₃COOH) in water. The K_a for acetic acid is 1.8 × 10⁻⁵. The pH of a 0.10 M acetic acid solution is 2.87. The pH of a 0.10 M acetic acid solution is 2.87. The pH of a 0.10 M acetic acid solution is 2.87.

Model 1 - Ion Concentrations for Acids and Bases

Acid	Concentration	Acid-Base Equilibrium	[H ⁺]	[OH ⁻]	pH or pOH
1. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
2. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
3. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
4. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
5. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
6. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
7. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
8. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
9. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0
10. Hydrochloric acid	0.10 M	CH ₃ COOH ⇌ CH ₃ COO ⁻ + H ⁺	0.10 M	1.0 × 10 ⁻¹⁴ M	1.0

1. What are the concentrations of H⁺ and OH⁻ in this solution?

2. Is this solution acidic or basic? Explain your answer.

3. What is the concentration of H⁺ in this solution? Explain your answer.

4. What is the concentration of OH⁻ in this solution? Explain your answer.

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