

ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example: $^{15}_{7}\text{N}^{+3}$

mass number ↓
atomic number ↑

charge ←

Always write the magnitude of the charge followed by the sign: N3+

7 protons
8 neutrons (15 - 7)
4 electrons

Can you find the other mistakes on this page?

Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
$^{12}_6\text{C}$						
$^7_3\text{Li}^+$						
$^{35}_{17}\text{Cl}^-$						
$^{39}_{19}\text{K}$						
$^{24}_{12}\text{Mg}^{2+}$						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

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to the earth. Droplets of water freeze around ice crystals as they bounce up and down within a storm cloud. Fall to earth when they get heavy. Vapor that changes directly into crystalline flakes because of freezing temperatures. Water vapor that forms droplets and falls to the earth. Large amount of droplets falling to the earth. Symbols rain O O

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